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(Solution) | Alcohols Phenols And Ethers |
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32.Chemistry | Basic concepts of chemistry
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11. molar mass = mass/moles; mass = 45.0 g;
find moles; $\Delta T = 90.2^{\circ}\text{C} - 80.1^{\circ}\text{C} = 10.1^{\circ}\text{C}$
 $\text{moles} = \frac{\Delta T_f}{K_f} \times \text{mass}$
 $= \frac{10.1}{5.12} \times 0.0750 = 0.147$
 $\text{moles} = 0.147$

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 $6.09 \text{ kJ} = -60.4 \text{ kJ}$ Practice Exercise The
total enthalpy change is the sum of the
changes of the individual steps: Check: The
components of the total energy change are
reasonable in...

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Answer/Explanation. Answer: Explanation: (i) (d) (ii) (e) (iii) (b) (iv) (a) In cleavage of ethers, higher alcohol and lower alkyl halide is formed except in tert. butyl methyl ether, in which $(\text{CH}_3)_3\text{C}^+$ is more stable so $(\text{CH}_3)_3\text{C-I}$ and CH_3OH is formed. Anisole with HI gives phenol and CH_3I because phenoxide ion is stabilised by resonance.

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