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SECTION 3 PROBLEMS Write the  
answer on the line to the left. Show all

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your work in the space provided. 1. 88%

The actual yield of a reaction is 22 g and the theoretical yield is 25 g. Calculate the percentage yield. 2. 6.0 mol of N<sub>2</sub> are mixed with 12.0 mol of H<sub>2</sub> according to the following equation:  $\text{N}_2(\text{g}) + 3\text{H}_2(\text{g}) \rightarrow 2\text{NH}_3(\text{g})$

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Stoichiometry. SECTION 1. SHORT

ANSWER Answer the following questions in the space provided. 1. \_\_\_\_\_ The coefficients in a chemical equation represent the (a) masses in grams of all reactants and products. (b) relative number of moles of reactants and products. (c)

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number of atoms of each element in each compound in a reaction.

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1. The following equation represents a

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laboratory preparation for oxygen gas:  
 $2\text{KClO}_3(\text{s}) \rightarrow 2\text{KCl}(\text{s}) + 3\text{O}_2(\text{g})$  ... CHAPTER 9 REVIEW ...

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