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The three most common gases given off in reactions are: • Oxygen

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- Hydrogen • Carbon Dioxide All these gases are colorless and odourless. There are tests we can do to establish which gas is given off in an experiment. This resource is designed for UK teachers. View US version .

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The van der Waals equation is $(P + \frac{a}{V^2})(V - nb) = nRT$ where P = pressure V = volume a = pressure correction constant unique to the gas b = volume correction constant unique to the gas n = the number of moles of gas T = absolute temperature The van der Waals equation includes a pressure and volume correction to take into account the interactions between molecules. Unlike ideal gases, the individual particles of a real gas have interactions with each

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other and have definite volume.

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The activity sets some critical thinking and pattern spotting tasks in the context of the noble gases. This can be used to develop skills in determining mathematical relationships between variables from graphs.

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