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1. Aqueous solutions of acids have a sour taste.
2. Acids change the color of acid-base indicators.
3. Some acids react with active metals and release hydrogen gas, H_2 .
4. Acids react with bases to

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~~Acids~~ produce salts and water. 5. Acids conduct electric ...

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3 is the Lewis acid because it accepts an
electron pair in forming a covalent bond.
It is not a Brønsted-Lowry acid, because

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~~Acids~~ it is not donating a proton. e. Which of the reactants is the Lewis base? Explain your answer. NF_3 is the Lewis base because it donates an electron pair in forming a covalent bond. 120 ACIDS AND BASES MODERN CHEMISTRY

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CHAPTER 14 REVIEW Acids and Bases

SECTION 2 SHORT ANSWER Answer

the following questions in the space

provided. 1. a. Write the two equations

that show the two-stage ionization of

sulfurous acid in water. _____

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CHAPTER 14 REVIEW . Acids and
Bases. SHORT ANSWER Answer the
following questions in the space provided.

1. a. Write the two equations that show the
two-stage ionization of sulfurous acid in
water. stage 1: $\text{H}_2\text{SO}_3(\text{aq}) + \text{H}_2\text{O}(\text{l}) \rightleftharpoons$

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 $\text{H}_3\text{O}^+(\text{aq}) + \text{HSO}_4^-(\text{aq})$ b. Which stage of ionization usually produces more ions? Explain your answer.

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binary acid. An acid that contains only 2 different elements: hydrogen and one

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Acids

more electronegative elements. oxyacid.
An acid that is a compound of hydrogen, oxygen, and a third element, usually a nonmetal. Arrhenius acid. A chemical compound that increases the concentration of hydrogen ions, H^+ , in aqueous solution.

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an acid that is a compound of hydrogen, oxygen, and a third element, usually a nonmetal. arrhenius acid. a substance that increases the concentration of hydronium ions in aqueous solution. strong acid. an acid that ionizes completely in a solvent.

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weak acid.

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CHAPTER 14 ACIDS AND BASES 5

When H_3PO_4 is added to water, the three acids that are present are H_3PO_4 , H_2PO_4^- , and HPO_4^{2-} . H_3PO_4 , with the largest K_a value, is the strongest of these

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